86. Calculate the amount of sulfur pyrite (kg) required to produce 100m^3 of roasting gas with a SO_2 concentration of 15% by reaction: $4FeS_2 + 11O_2 = 2Fe_2O3 + 8SO_2$:

A) 75.1;
B) 65.5;
C) 40.1;
D) 120.3;
E) 145.5.

87. Calculate the amount of pyrite (kg) containing 45% S needed to produce 1 ton of H_2SO_4:

A) 150.3;
B) 126.5;
C) 246.6;
D) 518.3;
E) 725.6.

88. The concentration of sulfuric acid obtained by double adsorption and double contacting from pyrite firing gas is equal to:

A) 96%;
B) 75%;
C) 98.3 %;
D) 73%;
E) 92.5%.

89. The process of firing sulfur pyrite is carried out in accordance with the following technological scheme:

A) complex cyclical;
B) cyclical;
C) open circuit;
D) bypass;
E) parallel.

90. Sulfuric acid production is carried out in accordance with the following scheme:

A) parallel;
B) cyclical;
C) open;
D) bypass;
E) combined.

91. To calculate the expenditure coefficient of sulfur pyrite by the reaction: $4FeS_2+11O_2=2Fe_2O_3+8SO_2$, - if it contains 35% sulfur:

A) 1.5340;
B) 1.6672;
C) 2.5158;
D) 1.4286;
E) 1.2324.

92. The optimal SO_2 content in the firing gas supplied to contact oxidation is:

A) 3%;
B) 15%;
C) 20%;
D) 12%;